

WJEC (Eduqas) Chemistry A-level

Physical and Inorganic 5.1 - Equilibrium Constants Flashcards

This work by [PMT Education](https://www.pmt.education) is licensed under [CC BY-NC-ND 4.0](https://creativecommons.org/licenses/by-nc-nd/4.0/)



Define dynamic equilibrium



Define dynamic equilibrium

Dynamic equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the backwards reaction. At equilibrium, the concentrations of reactants and products stay constant.



State le Chatelier's principle



State Le Chatelier's principle

When a reaction at equilibrium is subjected to a change in concentration, temperature or pressure, the position of equilibrium will move to counteract the change.



Explain the difference between K_p and K_c



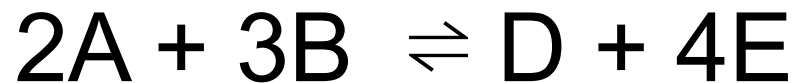
Explain the difference between K_p and K_c

K_c and K_p are both equilibrium constants.

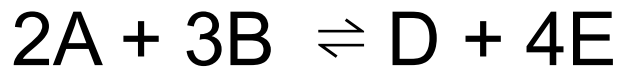
K_c is found using molar concentrations in the calculation whereas K_p uses partial pressures. If you are dealing with a reaction in which all the substances are gaseous it is generally easier to use K_p .



For the reaction below, deduce an equation for K_c :



For the reaction below, deduce an equation for K_c :



$$K_c = \frac{[D][E]^4}{[A]^2[B]^3}$$



Deduce the units for K_c where

$$K_c = \frac{[D][E]^4}{[A]^2[B]^3}$$



Deduce the units for K_c where $K_c = \frac{[D][E]^4}{[A]^2[B]^3}$

Replace the reactants/products in the K_c calculation with their units. Cancel any common units from the top and bottom of the fraction to find the units for K_c :

$$K_c = [(\text{mol dm}^{-3})(\text{mol dm}^{-3})^4] / [(\text{mol dm}^{-3})^2(\text{mol dm}^{-3})^3]$$

All the units cancel so K_c has **no units**.



Deduce the units for K_c where

$$K_c = \frac{[D][E]^3}{[A]^2[B]^3}$$



Deduce the units for K_c where $K_c = \frac{[D][E]^3}{[A]^2[B]^3}$

Replace the reactants/products in the K_c calculation with their units. Cancel any common units from the top and bottom of the fraction to find the units for K_c :

$$K_c = \frac{[(\text{mol dm}^{-3})(\text{mol dm}^{-3})^3]}{[(\text{mol dm}^{-3})^2(\text{mol dm}^{-3})^3]}$$
$$= (\text{mol dm}^{-3})^{-1} = \text{mol}^{-1} \text{dm}^3$$



Why does K_c change when the temperature of a reversible reaction in a closed system is changed?



Why does K_c change when the temperature of a reversible reaction in a closed system is changed?

The K_c value is only valid for a certain temperature. When the temperature changes, the position of equilibrium shifts and so the equilibrium concentrations of the products and reactants changes. This leads to a change in K_c .



Consider a reversible reaction where the forward reaction is endothermic. How will increasing the temperature affect K_c ?



Consider a reversible reaction where the forward reaction is endothermic. How will increasing the temperature affect K_c ?

If the temperature is increased then the forward endothermic reaction will be favoured so the position of equilibrium will move towards the products. This means the concentration of products will increase and concentrations of reactants will decrease. This leads to an increase in K_c as in the calculation for K_c the number on the top of the fraction will be larger.



Consider a reversible reaction where the forward reaction is exothermic. How will increasing the temperature affect K_c ?

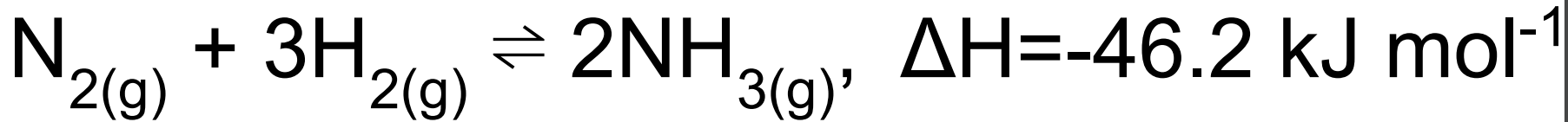


Consider a reversible reaction where the forward reaction is exothermic. How will increasing the temperature affect K_c ?

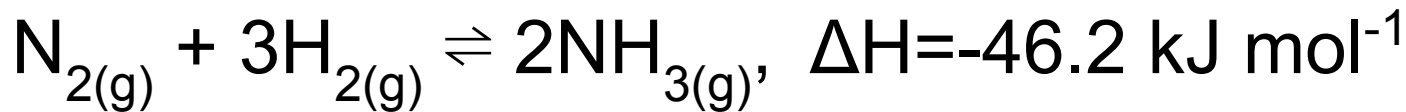
If the temperature is increased then the backwards endothermic reaction will be favoured so the position of equilibrium will move towards the reactants. This means the concentration of reactants will increase and concentrations of products will decrease. This leads to a decrease in K_c as in the calculation for K_c the number on the bottom of the fraction will be larger.



Explain why a compromised temperature of 450°C is used in the Haber process:



Explain why a compromised temperature of 450°C is used in the Haber process:



Since ΔH is negative we can deduce that the forward reaction is exothermic. Therefore a low temperature is required to shift equilibrium towards the products, and increase the yield of product. However the temperature cannot be too low as this will lead to a very slow rate of reaction. Therefore a compromised temperature is used so that the forward reaction is favoured but the rate of reaction is also relatively fast.



How does changing the concentration of a reactant or product affect the value of K_c ?



How does changing the concentration of a reactant or product affect the value of K_c ?

Changing the concentration of a reactant or product has no effect on K_c .



How does a catalyst affect the value of K_c ?



How does a catalyst affect the value of K_c ?

A catalyst has no effect on K_c . The catalyst will speed up the forward and backwards reactions at the same rate so the ratio of products to reactants will remain the same.



Consider a reversible reaction where the backwards reaction is exothermic. How could you increase the value of K_c ?



Consider a reversible reaction where the backwards reaction is exothermic. How could you increase the value of K_c ?

K_c will increase if the concentration of products increases. The forward reaction will be endothermic and so to favour this direction, and hence increase the concentration of products, the temperature needs to be increased.



Define partial pressure and total pressure in relation to a mixture of gases



Define partial pressure and total pressure in relation to a mixture of gases

Partial pressure - the pressure exerted by an individual gas in a mixture.

Total pressure - the sum of all the partial pressures of the individual gases in the mixture.



Define mole fraction



Define mole fraction

The proportion of a gas mixture that is made up of a particular gas.



Give the formula to calculate the mole fraction



Give the formula to calculate the mole fraction

Mole fraction =

number of moles of gas

total number of moles of gas in the mixture



Give the formula to calculate partial pressure of a gas



Give the formula to calculate partial pressure of a gas

Partial pressure of a gas =

(mole fraction of a gas) x (total pressure of the mixture)



What are the units of partial pressure?

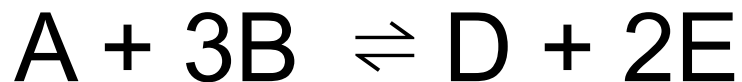


What are the units of partial pressure?

kPa



For the following gaseous reaction,
deduce an equation for K_p :



For the following gaseous reaction, deduce an equation for K_p : $A + 3B \rightleftharpoons D + 2E$

$$K_p = \frac{(p_D)(p_E)^2}{(p_A)(p_B)^3}$$



Deduce the units for K_p where

$$K_p = \frac{(pD)(pE)^2}{(pA)(pB)^3}$$



Deduce the units for K_p where $K_p = \frac{(p_D)(p_E)^2}{(p_A)(p_B)^3}$

$$K_p = [(kPa)(kPa)^2] / [(kPa)(kPa)^3] = kPa^{-1}$$

The units for K_p are kPa^{-1} .



Why does K_p change when the temperature of a closed system reversible reaction is changed?



Why does K_p change when the temperature of a closed system reversible reaction is changed?

The K_p value is only valid for a certain temperature. When the temperature changes, the position of equilibrium shifts which changes the mole fractions of the gases present. This affects their partial pressures and so K_p changes.



Consider a reversible reaction where the forward reaction is exothermic. How will decreasing the temperature affect K_p ?

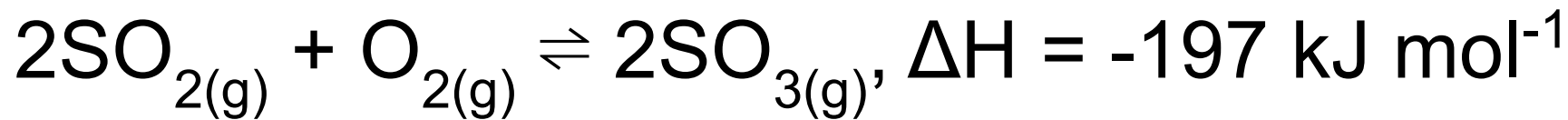


Consider a reversible reaction where the forward reaction is exothermic. How will decreasing the temperature affect K_p ?

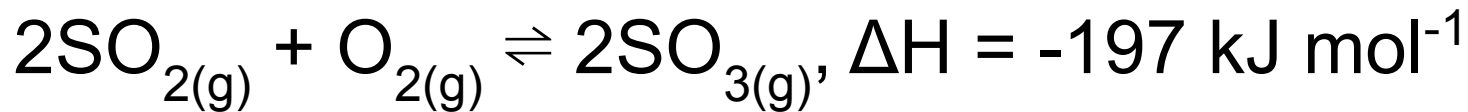
If the temperature is decreased then the forward exothermic reaction will be favoured so the position of equilibrium will move towards the products. This means the molar fraction and partial pressure of the products will increase and the molar fraction and partial pressure of the reactants will decrease. This leads to an increase in K_p as in the calculation for K_c the number on the top of the fraction will be larger.



How will K_p be affected if the temperature is increased in the following reaction?



How will K_p be affected if the temperature is increased in the following reaction?



The forward reaction is exothermic. If the temperature is increased then the equilibrium shifts to the left (in the endothermic direction) to counteract the change. This means that less product is formed so the partial pressure of SO_3 will decrease while the partial pressure of SO_2 and O_2 will increase. This will cause K_p to decrease.



What is the significance of the magnitude of an equilibrium constant?



What is the significance of the magnitude of an equilibrium constant?

The magnitude of the equilibrium constant indicates where the position of equilibrium lies. The larger the value of K_p or K_c , the further towards the right (towards the products) the position of equilibrium lies.

